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format with separate answer keys'

'Chemical Equations Gizmo ExploreLearning

December 28th, 2019 - Practice balancing chemical equations by changing the coefficients of reactants and products As the equation is manipulated the amount of each element is shown as individual atoms histograms or numerically Molar masses of reactants and products can also be calculated and balanced to demonstrate conservation of mass'

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Which set of diagrams shows a'

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December 28th, 2019 - While I am walking around I have the answer
key displayed on the document camera so students can check their
work for accuracy and develop any questions about counting atoms
After a few minutes I switch to the second page If students have no

questions we move on to the ExploreLearning Chemical Equations Prior Knowledge'

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classify five types of chemical reactions synthesis decomposition single
replacement double replacement and combustion While balancing the
reactions the number of atoms on each side is presented as visual
histogram and numerical data"***How to Balance Equations Printable
Worksheets**

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matter of making certain mass and charge are balanced on the
reactants and products side of the reaction arrow This is a collection
of printable worksheets to practice balancing equations The printable
worksheets are provided in pdf format with separate answer keys'**

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December 23rd, 2019 - equations Get the Gizmo ready Erase the chemical formulas in each text box Introduction In a chemical reaction the reactants are the substances that enter into the reaction and the products are the substances that are made in the reaction A

chemical reaction is balanced if the numbers of reactant atoms match the numbers of product atoms'

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 $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
 $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
 $3 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$
 $4 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$
 $5 \text{ Pb} + 2 \text{ OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$
 $6 \text{ Al} + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4$
 $3 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
 $8 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
 $9 \text{ C}_8\text{H}_{18}$

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